

Mark Scheme (Results)

Summer 2015

IAL Chemistry (WCH03/01)

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General Marking Guidance

- All candidates must receive the same treatment. Examiners must mark the first candidate in exactly the same way as they mark the last.
- Mark schemes should be applied positively. Candidates must be rewarded for what they have shown they can do rather than penalised for omissions.
- Examiners should mark according to the mark scheme not according to their perception of where the grade boundaries may lie.
- There is no ceiling on achievement. All marks on the mark scheme should be used appropriately.
- All the marks on the mark scheme are designed to be awarded. Examiners should always award full marks if deserved, i.e. if the answer matches the mark scheme. Examiners should also be prepared to award zero marks if the candidate's response is not worthy of credit according to the mark scheme.
- Where some judgement is required, mark schemes will provide the principles by which marks will be awarded and exemplification may be limited.
- When examiners are in doubt regarding the application of the mark scheme to a candidate's response, the team leader must be consulted.
- Crossed out work should be marked UNLESS the candidate has replaced it with an alternative response.
- Mark schemes will indicate within the table where, and which strands of QWC, are being assessed. The strands are as follows:
 - i) ensure that text is legible and that spelling, punctuation and grammar are accurate so that meaning is clear
 - ii) select and use a form and style of writing appropriate to purpose and to complex subject matter
 - iii) organise information clearly and coherently, using specialist vocabulary when appropriate

Question Number	Acceptable Answers	Reject	Mark
1(a)(i)	Sodium (ion)/ Na^+ If name AND formula are given BOTH must be correct	Na	1

Question Number	Acceptable Answers	Reject	Mark
1(a)(ii)	Carbonate (ion)/ CO_3^{2-} OR CO_3^{-2} / CO_3^{--} OR Hydrogencarbonate (ion)/ HCO_3^- ALLOW Hydrogen carbonate (ion) If name AND and formula are given BOTH must be correct		1

Question Number	Acceptable Answers	Reject	Mark
1(a)(iii)	$\text{Ca(OH)}_2(\text{aq}) + \text{CO}_2(\text{g}) \rightarrow \text{CaCO}_3(\text{s}) + \text{H}_2\text{O}(\text{l})$ Reactants with state symbols (1) Products and state symbols (1) Allow All formulae correct but one or more errors in state symbols (1) All formulae and state symbols correct but incorrect balancing numbers included (1) $\text{Ca}^{2+}(\text{aq}) + 2\text{OH}^-(\text{aq})$ for $\text{Ca(OH)}_2(\text{aq})$ Ignore multiples if equation is balanced	$\text{H}_2\text{O}(\text{aq})$ $\text{Ca(OH)}_2(\text{l})$ $\text{Ca}^{2+}(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{CaCO}_3(\text{s})$	2

Question Number	Acceptable Answers	Reject	Mark
1(b)(i)	Crimson/red/ dark red/bright red / persistent red/scarlet (coloured flame)	Orange Brick red Carmine	1

Question Number	Acceptable Answers	Reject	Mark
1(b)(ii)	White precipitate ALLOW White solid/crystals ppt/ppte for precipitate Ignore comments about getting darker/turning purple on standing	Cream ppt Yellow ppt Off-white ppt Additional incorrect observations eg white ppt and effervescence or steamy fumes Change on standing to cream or yellow	1

Question Number	Acceptable Answers	Reject	Mark
1(b)(iii)	TEST Add dilute (aqueous) ammonia (solution) / NH_3 / NH_3 (aq) ALLOW Dilute NH_4OH / ammonium hydroxide IGNORE Additional test with concentrated NH_3 (1) RESULT Precipitate dissolves / (colourless) solution forms ALLOW mixture dissolves / precipitate disappears/ solid dissolves / precipitate is soluble (1) Second mark depends on use of ammonia in first, even concentrated. ALLOW TEST add concentrated sulfuric acid to ppt (1) RESULT Steamy fumes (only)/ no brown AND no purple fumes ALLOW White fumes (1) Second mark depends on use of sulfuric acid.	Just "ammonia / NH_3" Use of ammonia on glass rod Incorrectly identified precipitate dissolves e.g. strontium chloride dissolves White smoke	2

Question Number	Acceptable Answers	Reject	Mark
1(b)(iv)	(goes) dark / purple / grey ALLOW Black / lilac (1) Silver / Ag (forms) (1)	Goes blue-black Silver colour/ mirror Ag ⁺ / silver ions	2

Question Number	Acceptable Answers	Reject	Mark
1(c)	$\text{Sr}^{2+} + \text{CO}_3^{2-} \rightarrow \text{SrCO}_3$ Ignore state symbols, even if incorrect Ignore full equation, written as "rough" work and mark ionic equation only.		1

Total for Question 1 =11 marks

Question Number	Acceptable Answers	Reject	Mark
2(a)	TEST Add PCl_5 / phosphorus(V) chloride / phosphorus pentachloride / SOCl_2 / thionyl chloride / sulphur dichloride oxide (1) RESULT Mark depends on correct reagent, but allow PCl_5 (aq) Steamy / misty / white fumes ALLOW Gas for fumes (1) Ignore incorrect identification of fumes OR TEST Add sodium / Na (1) RESULT Mark depends on correct reagent Effervescence / bubbling / fizzing Ignore incorrect identification of fumes and tests for products white solid (forms) / sodium dissolves mixture gets hot (1)	Acidified PCl_5 / PCl_5 (aq) Acidified dichromate(VI) PCl_3 Test to form an ester Any smoke Just "HCl fumes" Just "gas turns litmus red" Just "hydrogen"	2

Question Number	Acceptable Answers	Reject	Mark
2(b)	(primary / secondary / tertiary) Alcohol and carboxylic acid ALLOW ROH and RCOOH R_2CHOH/ R_3COH for ROH $C_nH_{2n+1}OH$ for ROH RCO_2H for RCOOH Phenol(s) (as one alternative) Fatty acid / alkanoic acid for carboxylic acid	diol carboxyl cyclic alcohol specific alcohol eg ethanol	1

Question Number	Acceptable Answers	Reject	Mark
2(c)	Z identified as tertiary alcohol (1) Justification: Any one from Test with litmus Not (carboxylic) acid because there is no change (in (blue) litmus paper) It's an alcohol because there is no change (in (red / blue) litmus paper) It is neutral /not an acid or an alkali because there is no change (in (red / blue) litmus paper) Test with dichromate It is a tertiary alcohol because it can't be oxidized (by acidified dichromate(VI))/ doesn't react with acidified dichromate(VI) It is not a primary or secondary alcohol because it can't be oxidized (by acidified dichromate(VI))/ doesn't react with acidified dichromate(VI) IGNORE Not an amine (1) If more than one justification is given, both must be correct		2

Question Number	Acceptable Answers	Reject	Mark
2(e)(i)	Molecular ions have same m/e ALLOW same molecular ion isomers have same molar mass / molecular mass molecular ion with same mass same maximum m/e value same peak furthest to right same last peak Parent ion / M^+ for molecular ion IGNORE Reference to peak heights	Same fragments Same m/e value for highest peak Similar for "same"	1

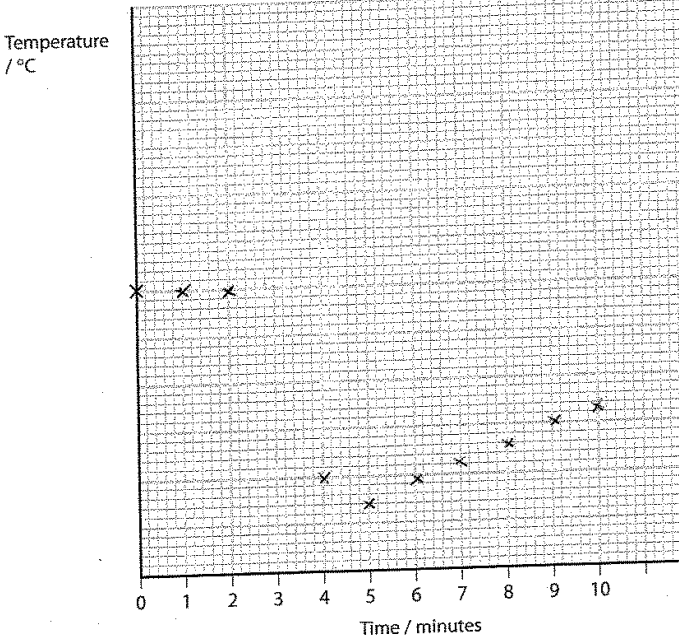
Question Number	Acceptable Answers	Reject	Mark
2(e)(ii)	They both have an (absorption) peak for (wavenumber of) alcohol / hydroxyl group / O-H ALLOW both have peak for -OH / OH frequency / wavelength for wavenumber IGNORE wavenumber values have peak with specific shape for OH	Absorption for C-OH	1

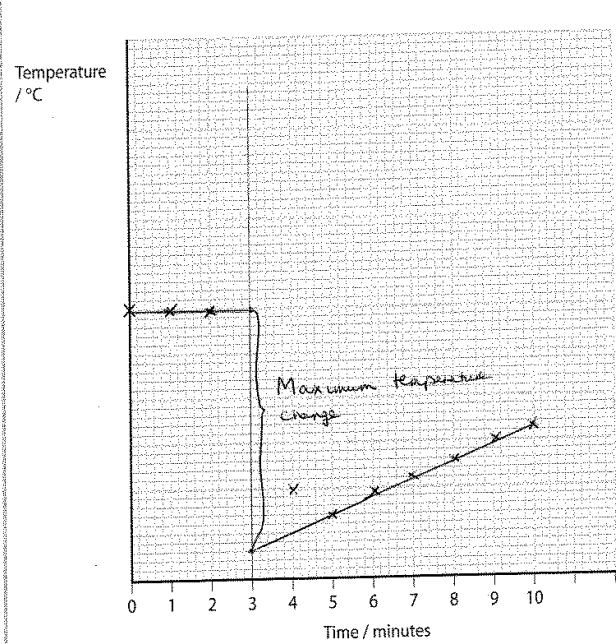
Total for Question 2 = 10 marks

Question Number	Acceptable Answers	Reject	Mark
3(a)(i)	Correct final answer with + sign, 3 sf and units scores 3 $(25 \times 4.18 \times 10.5) = 1097.25 \text{ (J) / 1.097 kJ}$ Ignore sign if given (1) $\text{Mol NH}_4\text{Cl} = (5.00/53.5) = 0.09346/0.0935$ (1) $\Delta H_{\text{solution}} = (+1.097/0.09346)$ $(= +11.7376 / +11.7406)$ = +11.7 kJ mol⁻¹ OR +11700 J mol⁻¹ Sign, unit and sf must be correct for third mark Use of 2sf earlier may lead to an inaccurate answer (1) ALLOW Final answer = +11.8 kJ mol⁻¹ from rounding of MP1 and/or MP2 (3) TE from each step to the next If mass used is 30 g Energy transferred = 1316.7 J $\Delta H_{\text{solution}} = +14.1 \text{ kJ mol}^{-1}$ max (2) If mass used is 5 g Energy transferred = 219.45 J $\Delta H_{\text{solution}} = +2.35 \text{ kJ mol}^{-1}$ max (2)	Answers not to 3 sf No sign or negative sign	3

Question Number	Acceptable Answers	Reject	Mark
3(a)(ii)	First mark is for calculation of error. Second mark is for comparison of temperature error to mass error. Uncertainty in mass = $(0.005 \times 100 \times 2/5.00) = (\pm)0.2\%$ (1) Uncertainty / error in mass measurement (much) smaller than uncertainty in temperature reading (1) Second mark depends on first being correct, but allow second mark if mass error is 0.1% (as 0.005 not doubled)	Just "0.2% is negligible / very small"	2

Question Number	Acceptable Answers	Reject	Mark
3(b)(i)	Points (close to the) horizontal from starting temperature at 0, 1 and 2 (and 3) minutes (1) Points (on a line) rising from a minimum up to 10 minutes (at least 2 points needed at the warming up stage for extrapolation.) The minimum can be at 4, 5, 6, 7 or 8 minutes. (1)	Large change of temperature at 3 minutes Cooling curve instead of warming curve	2



Question Number	Acceptable Answers	Reject	Mark
3(b)(ii)	Line through temperature points where warming occurs extrapolated back to 3 minutes. ALLOW Line at minimum temperature shown as staying horizontal and extrapolated back (1) Max temperature change indicated as vertical difference between starting temperature and extrapolated line at 3 minute (1) TE if cooling curve drawn in 3(b)(i) for both marks. 		2

Question Number	Acceptable Answers	Reject	Mark
3(b)(iii)	To check water temperature is steady / constant OR To deduce temperature at 3 mins / at start by extrapolation of line ALLOW to allow water temperature to equilibrate with surroundings/ to reach temperature of surroundings/ to acclimatise IGNORE to get initial temperature accurate	Water temperature may change Minerals in water may affect result	1

Question Number	Acceptable Answers	Reject	Mark
3(c)(i)	Heat must be supplied (and cannot be measured) ALLOW impossible to tell when/if reaction is complete reaction goes to equilibrium/ is reversible IGNORE reference to gases escaping / products are gases / hazards	Just " because it is endothermic" Needs high temperature	1

Question Number	Acceptable Answers	Reject	Mark
3(c)(ii)	$\text{NH}_4\text{Cl(s)}$ $\xrightarrow{\Delta H_{\text{reaction}}}$ $\text{NH}_3\text{(g)} + \text{HCl(g)}$ $\Delta H_1 \downarrow$ $\Delta H_2 + \Delta H_3 \downarrow$ $\text{NH}_4\text{Cl(aq)}$ $\xleftarrow{\Delta H_4}$ $\text{NH}_3\text{(aq)} + \text{HCl(aq)}$ OR 2 separate parallel arrows for $\Delta H_2 + \Delta H_3$ OR $\Delta H_2 \Delta H_3$ next to one arrow without being separated by + ALLOW Arrows reversed if signs of enthalpy changes are reversed. IGNORE Any water molecules added/ aq signs / other reactant species Arrow size		1

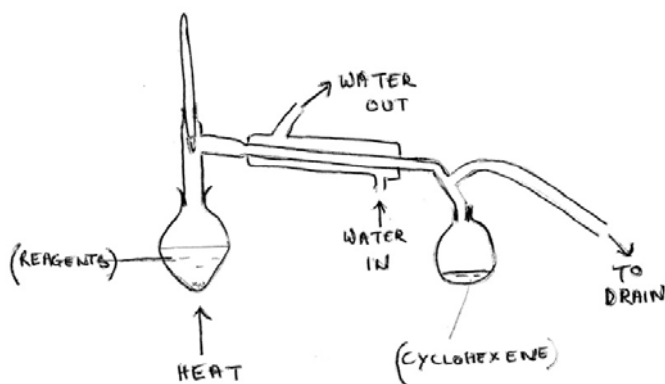
Question Number	Acceptable Answers	Reject	Mark
3(c)(iii)	$\Delta H_{\text{reaction}} = \Delta H_1 - \Delta H_2 - \Delta H_3 - \Delta H_4$ ALLOW any order of terms with correct signs Any correct use of brackets No TE on incorrect cycle		1

Total for Question 3 =13 marks

Question Number	Acceptable Answers	Reject	Mark
4(a)	Two different hazards must be given to score 2 marks. Phosphoric acid corrosive ALLOW burns skin/ damages skin (1) Cyclohexanol / cyclohexene (in)flammable ALLOW Irritant (1) IGNORE Comments on glass wool, calcium chloride Cyclohexene / cyclohexanol is volatile	Additional hazards e.g. irritant harms skin carcinogenic Additional hazards e.g. explosive carcinogenic	2

Question Number	Acceptable Answers	Reject	Mark
4(b)	Correct final answer scores (2) Mass of 12 cm³ C₆H₁₁OH = 12 x 0.962 = 11.544 / 11.54 / 11.5 (g) (1) Number of moles = (11.544 / 100 = 0.11544) = 0.115 / 0.12 (mol) ALLOW TE from incorrect mass (1) Ignore sf except 1 sf	0.11	2

Question Number	Acceptable Answers	Reject	Mark
4(c)	Flask with heat source AND stillhead AND a closed system to the left hand side of the outlet to the condenser. Heat source can be electrical heater, water bath ALLOW bunsen or just arrow ALLOW appropriate tubing or flask with long neck as alternative to stillhead (1) Bulb of thermometer opposite opening to condenser (1) Water condenser sloping downwards AND direction of water (1) Connected to receiver with a vent OR delivery tube to an open narrow necked flask (1) Ignore fractionating column if included. Drawing showing reflux distillation scores max 1 for water direction in condenser.	Conical flask Sealed receiver, beaker	4

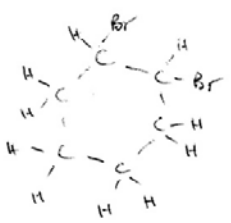


Question Number	Acceptable Answers	Reject	Mark
4(d)	Dehydrating agent removes water in a (chemical) reaction OR causes two H and one O atoms to be lost (in a reaction) OR removes the elements of water (from reactant molecules) OR removes water from molecules of a compound ALLOW answers indicating a reaction occurs eg H^+ protonates OH in alcohol forming water removes water causing bonds to break reference to elimination reactions (1) Drying agent removes water mixed with other materials OR removes water from a mixture OR removes water in a physical change ALLOW Absorbs water (1) "A dehydrating agent removes water in a reaction but there is no reaction when a drying agent removes water" scores 1	Reference to removal of solvents other than water	2

Question Number	Acceptable Answers	Reject	Mark
4(e)	Glass wool less absorbent OR No cyclohexene left on wool OR filtration is faster through glass wool OR filter paper absorbs liquids/ product/ mixture IGNORE yield is higher with glass wool/ lower with filter paper more efficient filtration		1

Question Number	Acceptable Answers	Reject	Mark
4(f)	Look at final answer. If correct award 3 marks. There are several correct methods. All involve calculating a number of moles of cyclohexene, a mass of cyclohexanol and the use of the 75% but these stages can be done in different orders. EITHER Need theoretical yield of $(10.0 \times 100/75) = 13.3333 / 13.33 / 13.3 \text{ g}$ (1) $13.3333 \text{ g} = (13.3333/82) = 0.1626 / 0.163 \text{ mol cyclohexene}$ (1) $0.1626 \text{ mol cyclohexanol} = \mathbf{16.26 / 16.3 / 16 \text{ g}}$ (1) OR Mol of cyclohexene $= (10/82) = 0.12195$ (1) Mol of cyclohexanol $= (0.12195 \times 100/75) = 0.1626$ (1) Mass of cyclohexanol $= (0.1626 \times 100) = \mathbf{16.26 / 16.3 / 16 \text{ g}}$ (1) OR Mol of cyclohexene $= (10/82) = 0.12195$ (1) Theoretical mass of cyclohexanol $= (0.12195 \times 100) = 12.195/12.2 \text{ g}$ (1) Mass of cyclohexanol $= (12.2 \times 100/75) = \mathbf{16.26 / 16.3 / 16 \text{ g}}$ (1) ALLOW 16.2 (g) in all methods from rounding 9.146 (g) from incorrect use of 75% scores (2) Ignore SF in final answer except 1 SF	Theoretical yield $= (10.0 \times 75/100) = 7.5 \text{ g}$ $(0.12195 \times 75/100) = 0.09146$	3

Question Number	Acceptable Answers	Reject	Mark
4(g)(i)	Brown / red-brown / orange / yellow/ yellow-brown to colourless ALLOW Brown / red-brown / orange / yellow is decolorised. IGNORE Clear for colourless	Red to colourless	1

Question Number	Acceptable Answers	Reject	Mark
4(g)(ii)	 ALLOW Rings with CH ₂ and/or CHBr IGNORE Angles in ring Placing of H and Br inside or outside ring	Benzene ring Just skeletal formula/ molecular formula Bromoalcohols Non-adjacent Br atoms	1

Total for Question 4 = 16 marks

